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AW Rasmussen

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 18mol br₂ identify the limiting reactant
 s and excess reactant s the limiting
 reactant is o₂ since it would yield the
 least amount of product 18 mol br₂ the
 excess reactant is c₂h₃br₃ since its
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 up to 54 mol br₂ think about your

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 smaller amount of product is the
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 limiting reagent in this reaction step 5
 the reactant that produces a larger
 amount of product is the excess reagent
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react with 2 moles of oxygen which of
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at the end and how many moles of it
are leftover hydrogen 2 moles leftover
hydrogen 1 mole leftover

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c a o h 2 and 16 35 g of h₂so₄ h 2 s o 4
a determine the limiting reagent b
determine the number of moles of h₂o
h 2 o produced c determine the number
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determine the number of grams of
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