

# Canada Pr Points Calculator

JE Gale

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equations  
 $\text{C}_6\text{H}_{10}\text{O}_2 + 12\text{CO}_2 + 10\text{H}_2\text{O} \rightarrow 3\text{HBr} + 1\text{KHCO}_3 + 1\text{H}_2\text{O} + 1\text{KBr} + \text{CO}_2$   
 $\text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{NaBr} \rightarrow 5\text{H}_2\text{SO}_4 + 3\text{H}_2\text{O} + 3\text{Na}_2\text{SO}_4$   
 using the following equation

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 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$   
 $\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{NaBr} \rightarrow 5\text{H}_2\text{SO}_4 + 3\text{H}_2\text{O} + 3\text{Na}_2\text{SO}_4$

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web may 28 2020 write the balanced equation for the reaction and calculate

how much  $\text{SiO}_2$  is required to produce 300 kg of  $\text{SiC}$  answer  
 $\text{SiC} + 2\text{CO} \rightarrow \text{SiO}_2 + 2\text{C}$

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 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$   
 $\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{NaBr} \rightarrow 5\text{H}_2\text{SO}_4 + 3\text{H}_2\text{O} + 3\text{Na}_2\text{SO}_4$   
 solve the following stoichiometry problems  
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balanced equations make it possible to compare different elements through the stoichiometric factor discussed earlier this is the mole ratio between two factors in a chemical reaction found through the ratio of stoichiometric coefficients

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 $\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{NaBr} \rightarrow 5\text{H}_2\text{SO}_4 + 3\text{H}_2\text{O} + 3\text{Na}_2\text{SO}_4$

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 $\text{Mg} + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$

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coefficient of at least one stuck review  
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