

# Daisy Keech Meal Plan

**Kathleen Armour**

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by the total number of molecules of all the components in the solution

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0023 g mol 156 g kbr

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web calculate the mass percent mole fraction molality and molarity of hexane  $C_6H_{14}$  6 12 14 1 g mol 98 g mol 1 when it is mixed with heptane  $C_7H_{16}$  7 12 16 1 g mol 100 g mol 1 in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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