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## Karin Nielsen-Saines

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web the balanced chemical equation for this reaction is  $\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$  in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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web mar 4 2018  $\text{Mg}$   $2\text{HNO}_3$   $\text{Mg}$   $\text{NO}_3$   $2$   $\text{H}_2$  on the left side there are  $\text{Mg}$  1 magnesium in elemental form  $\text{H}_2$  there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens  $\text{N}_2$  there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens  $\text{O}_6$

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web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate  $\text{MgCO}_3$  and two moles of nitric acid  $\text{HNO}_3$  react to form one mole of magnesium nitrate  $\text{Mg}(\text{NO}_3)_2$  one mole of water  $\text{H}_2\text{O}$  and one mole of carbon dioxide  $\text{CO}_2$  show chemical structure image reactants

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product side balancing nitrogen by multiplying  $\text{hno}_3$  by 2 we get  $\text{mg}$  s magnesium 2  $\text{hno}_3$  l nitric acid  $\text{mg}$  no 3 2 s magnesium nitrate h 2 g hydrogen gas oxygen and hydrogen get automatically balanced by balancing nitrogen  
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