

The Moon And The Yew Tree

Michael Hoelscher

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4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3
 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times \frac{101.96 \text{ g Al}_2\text{O}_3}{302.3 \text{ g Al}_2\text{O}_3} = 26.98 \text{ g Al}_2\text{O}_3$
c determine the percent yield $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{26.0 \text{ g}}{26.98 \text{ g}} \times 100 = 86.01\%$

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