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 $2 \text{ CO}_2 \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \frac{5}{2} \text{ O}_2 \text{ g} \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O g} + \text{H}_2 \text{ g} + \frac{1}{2} \text{ O}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g} \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O g} \quad \Delta h = 283 \text{ kJ}$
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worksheet calculate h for the reaction c

2 h 4 g h 2 g c 2 h 6 g from the

following data c 2 h 4 g 3 o 2 g 2 co 2 g

2 h 2 o l h 1411 kj mole c 2 h 6 g 7 2 o

2 g 2 co 2 g 3 h 2 o l h 1560 kj mole h 2

g 1 2 o 2 g h 2 o l h 285 kj mole

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the enthalpy change of an overall

process is equal to the sum of the

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delta h for the target reaction ce 2 no2

g 1 2 o2 g n2o5 g given the following

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that regardless of the multiple stages

or steps of a reaction the total enthalpy

change for the reaction is the sum of all

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2019 which determines the enthalpy

change value hess s law is written as

$\Delta h_{\text{target}} = \Delta h_1 + \Delta h_2 + \Delta h_3$ or $\Delta h_{\text{target}} = \sum \Delta h_{\text{known}}$

Δh_{known}

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can be expressed as the sum of two or

more steps the enthalpy change for the

overall process is the sum of the Δh

values for each step to use hess s law

two principles must be understood one

if an equation is reversed the sign of

the Δh value is also reversed two if an

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directions for each of the following

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baron lily zmachinski introduction in

order to calculate the enthalpy change

for the combustion of magnesium oxide

mg s 1 2o2 g mgo s we used

additivity of heats of reaction hess s law

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enthalpy change for the reaction will be

the sum of all those changes no matter

how many different steps or stages in

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equations for the reactions in the

experiment done are as follows 1 naoh

s na aq oh aq

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is the sum of two or more other

reactions the Δh for the overall process

must be the sum of the Δh values of the

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Δh is independent of the path that a

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the solubility naoh in hcl and the

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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