

Obturator In Cleft Palate

N Colangelo

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identify the chemical equation the spectator ions $\text{Pb}^{2+} \text{aq}$ $\text{NO}_3^- \text{aq}$ $\text{PbI}_2 \text{s}$ $\text{KNO}_3 \text{aq}$

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$ $\rightarrow$ $2\text{Cr}^{3+} \text{aq}$ $7\text{H}_2\text{O}$ label 4 2 3

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 answer key to practice problems on net ionic equations 1
 $\text{Molecular: AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} \downarrow + \text{KNO}_3$
 $\text{Total Ionic: Ag}^+ + \text{NO}_3^- + \text{K}^+ + \text{Cl}^- \rightarrow \text{AgCl} \downarrow + \text{K}^+ + \text{NO}_3^-$
 $\text{Net Ionic: Ag}^+ + \text{Cl}^- \rightarrow \text{AgCl} \downarrow$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$
 the spectator ions in KOH and HNO_3 are K^+ and NO_3^- how to identify and list the

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 molecular equation showing the complete chemical formulas of reactions and products $\text{Pb(NO}_3)_2 + 2\text{KCl} \rightarrow \text{PbCl}_2 \downarrow + 2\text{KNO}_3$
 complete ionic equation an equation written with all soluble strong electrolytes shown as ions

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