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 $2 \text{ CO}_2 \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \frac{5}{2} \text{ O}_2 \text{ g}$ $\Delta h = 2 \text{ h}_2 \text{ g} + \text{C}_2\text{H}_6 \text{ g}$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O g} \rightarrow \text{H}_2 \text{ g} + \frac{1}{2} \text{ O}_2 \text{ g}$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2 \text{ O g}$ $\Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g}$

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 $\text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$ from the following data
 $\text{C}_2\text{H}_4 \text{ g} + 3 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2 \text{ O l}$ $\Delta h = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} + 7 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2 \text{ O l}$ $\Delta h = 1560 \text{ kJ}$
 $\text{H}_2 \text{ g} + \frac{1}{2} \text{ O}_2 \text{ g} \rightarrow \text{H}_2 \text{ O l}$ $\Delta h = 285 \text{ kJ}$
 mole

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changes of its individual steps hess s law example 12 1
determine ΔH for the target reaction $\text{C}_2\text{H}_4\text{O}_2 + 2\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$ given the following information

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"hess s law" 2019 which determines the enthalpy change value hess s law is written as $\Delta H_{\text{target}} = \Delta H_1 + \Delta H_2 + \Delta H_3$ or $\Delta H_{\text{target}} = \sum \Delta H_{\text{known}}$

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zmachinski introduction in order to calculate the enthalpy change for the combustion of magnesium oxide MgO we used

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cohen 2016 the equations for the reactions in the experiment done are as follows $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O h 55 8 kj as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}}$

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