

# If I Have To Be Haunted

J Dewey

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of  $\text{C}_3\text{H}_8$  what is the percent yield  $3 \text{CO} + 7 \text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3 \text{H}_2\text{O}$

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of co would be produced step one figure out how many moles of the limiting reagent you have o 2 16 2 32 grams mole 21 3 grams 0 66 moles 32 g m

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web limiting reactant and percent yield practice name 1 consider the following reaction  $\text{nh}_4\text{no}_3 + \text{na}_3\text{po}_4 \rightarrow \text{nh}_4\text{ } 3\text{po}_4$  nano3 which reactant is limiting assuming we started with 30.0 grams of

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web limiting reagents and percentage yield worksheet 1 consider the reaction  $\text{I}_2\text{O}_5 + 5 \text{CO} \rightarrow 5 \text{CO}_2 + \text{I}_2$  a 80.0 grams

of iodine v oxide  $\text{I}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide  $\text{CO}$  determine the mass of iodine  $\text{I}_2$  which could be produced 80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$   $\times$  1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$   
**7.3 limiting reactant and percent**

**yield problems** - Jun 13 2023  
web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3