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web oxidation reduction redox reactions oxidation $\text{Cu} \rightarrow \text{Cu}^{2+} + 2\text{e}^{-}$
 reduction $\text{MnO}_4^{-} + 4\text{H}^{+} + 3\text{e}^{-} \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$
 when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} : \text{MnO}_4^{-}$ is

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 $\text{NO}_2 + \text{Al} + \text{NH}_3 \rightarrow \text{AlO}_2 + \text{NH}_4^{+}$
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solution what is the coefficient in front of OH^{-} in the balanced equation
 previous topic next topic

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 $\text{MnO}_4^{-} + \text{H}_2\text{O} + 2\text{e}^{-} \rightarrow \text{MnO}_2 + 4\text{OH}^{-}$
 what will be the stoichiometric coefficient of H_2O if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution
 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$
 balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution
 $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{H}_2\text{O}_2 \rightarrow 2\text{HCN} + 2\text{H}_2\text{O}$
 make electrons equal convert to basic solution
 $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-} \rightarrow 4\text{HCN} + 2\text{H}_2\text{O}$
 multiplied by a factor of 2
 $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 + \text{H}_2\text{O}_2 \rightarrow 2\text{OH}^{-} + 3\text{H}_2\text{O}$
 the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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10 problem 1 $\text{Cr}_2\text{O}_7^{2-}$ Fe^{2+} Cr^{3+} Fe^{3+} solution 1 balanced half
reactions 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-}$ 2Cr^{3+} $7\text{H}_2\text{O}$ Fe^{2+} Fe^{3+} e 2 equalize the
electrons 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-}$ 2Cr^{3+} $7\text{H}_2\text{O}$ 6Fe^{2+} 6Fe^{3+} 6e^- multiplied
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web balance each half reaction in basic solution 4 5 6 7
 $\text{Cr}_2\text{O}_7^{2-}$ 2 no SO_4^{2-} 2 MnO_2 Cr^{3+} no3 SO_2 Mn_2O_3 balance each
redox reaction in acid solution using the half reaction
method 8 H_2O_2 $\text{Cr}_2\text{O}_7^{2-}$ 2 9 TeO_3 2 N_2O_4 O_2 Cr^{3+} te 10 ReO_4^- 11
 PbO_2 io i2 Pb^{2+} 12 as no3 io3 re io3 H_2AsO_4^-

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reactions balance the atoms in each half reaction add the
two half reactions together and cancel out common terms
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redox reactions the equation is separated into two half

equations one for oxidation and one for reduction the
equation is balanced by adjusting coefficients and adding H^+
 O^{2-} and e^- in this order balance the atoms in the equation
apart from O^{2-} and H^+

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web in the redox equation below that occurs in our stomach
which of the following pairs identifies the reducing and
oxidizing agents respectively c $6\text{H}_8\text{O}_6$ $2\text{H}_2\text{NO}_2$ c $6\text{H}_8\text{O}_6$
 $2\text{H}_2\text{O}$ 2NO choose 1 answer c $6\text{H}_8\text{O}_6$ and no 2a c $6\text{H}_8\text{O}_6$
and no 2no 2 and c $6\text{H}_8\text{O}_6$ b no 2 and c $6\text{H}_8\text{O}_6$ h and no
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