

# Men Of The Cross

## D Kirk

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59 mol of  $\text{Fe}_2\text{O}_3$  how many grams of  
 so 3 can react with it using the mole  
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 $\text{AgNO}_3$  1c 20 mol  $\text{H}_2\text{O}$  1d 10 mol no 2a  
 38 mol  $\text{N}_2\text{H}_4$  2b 19 mol  $\text{N}_2\text{O}_4$  2c 76 mol  
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 was produced how many moles of  $\text{O}_2$   
 were required 0 326 mol  $\text{O}_2$  b if  $4.0 \times 10^3$   
 moles of oxygen reacted how many  
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 produced through the following  
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 how many grams

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 and the mass in milligrams of diatomic  
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 $2\text{AgNO}_3 + \text{BaCl}_2 \rightarrow 2\text{AgCl} + \text{Ba(NO}_3)_2$  how many  
 grams of  $\text{AgCl}$  silver chloride are  
 produced from 5 0 g of  $\text{AgNO}_3$  silver  
 nitrate 5 how

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 g 13 8 text g 1 3 8 g 13 point 8 start  
 text g end text of sodium chloride  
 express the answer using 3 significant  
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