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 $^{137}_{55}\text{Cs} \rightarrow ^{137}_{56}\text{Ba} + x$ because the mass number is the same on both sides of the equation x must have a

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 $2\text{Li} + \text{N}_2 \rightarrow 2\text{Li}_3\text{N}$
 $\text{Ca} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $\text{Cu} + \text{HNO}_3 \rightarrow \text{Cu(NO}_3)_2 + \text{H}_2\text{O} + \text{NO}$
 balance the following equations
 $\text{Fe} + \text{Cl}_2 \rightarrow \text{FeCl}_3$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $\text{As}_2\text{O}_3 + \text{NaOH} \rightarrow \text{Na}_3\text{AsO}_3 + \text{H}_2\text{O}$

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 $^{10}_{0}\text{n} + ^{235}_{92}\text{U} \rightarrow ^{137}_{52}\text{Te} + ^{90}_{38}\text{Sr} + 2\text{He}$

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