

The Law Of One

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 $\text{a} \times 2 \text{e}^{-}$ reduction $\text{MnO}_4^{-} \rightarrow \text{Mn}^{2+}$
 $\text{a} \times 5 \text{e}^{-}$ when the half reactions listed above are combined into a
 balanced chemical equation the ratio of the coefficients $\text{Cu} \rightarrow \text{MnO}_4^{-}$ is

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solution 2cn au au cn 2 e 2e 2h o2 h2o2 2 make electrons equal convert to basic solution 4cn 2au 2au cn 2 2e multiplied by a factor of 2 2e 2h2o o2 h2o2 2oh 3 the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation

half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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