

# Violines Bajo El Agua

**O García**

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in order to find the molarity the mass of  
the butane should be divided by the  
moles after further calculations the  
molarity of the butane gas was

calculated to be 56 81 grams moles this  
lab is done with tons of sources of error  
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web this data shows that there was a  
gain the mass of the lighter which  
shows that there was a certain amount  
of butane 0 38 grams 0 02 grams gas  
that was successfully collected during  
the gas collection there is a mild  
influence of barometric pressure during  
both trials as shown above in the raw  
data

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determine the number of moles of  
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ideal gas and that one mole has a  
volume of 22 4 l at stp v 1 n 1 v 2 n 2 0  
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12 experimental molar mass g mole of  
butane 0 23 g 0 0040 moles 58 g mole  
13 the molecular formula of butane is c  
4 h 10

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moles and mass will be required to  
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the experiment the mass of a butane  
lighter was measured before using the  
butane gas to deplete the water to the  
80ml mark on a graduated cylinder  
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 arbrisseaux qui associait l étendue et la  
 rigueur de l information à une forme  
 accessible à un large public oeuvre de  
 botanique et d ethnobotanique mais  
 aussi tentative d écriture était vite  
 apparu comme la plus riche synthèse  
 de langue française en son domaine