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of solutions of acids and bases through
the titration process we are able to
identify physical changes to the mixture
such as the colour change to indicate
the end point of the experiment

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known acidic or basic substance
through acid base reactions the analyte
titrand is the solution with an unknown
molarity the reagent titrant is the
solution with a known molarity that will
react with the analyte

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experiment was to identify the
unknown solid acid 17 by calculating its
molar mass this was done by titration of
the acid with sodium hydroxide in order
to find the volume of sodium hydroxide
required to reach the end point and

that value was used in unit analysis to
find the molar mass and identify the
unknown solid acid

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web by utilizing the half equivalence
point on the titration curve the pka of
the acid analyte was determined to be 3
655 furthermore using the
stoichiometric ratio between the naoh
and the acid analyte 1 1 we could
experimentally determine the number
of moles of the acid in the solution to be
000731 moles making the experimental
molar mass

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titration abstract the purpose of this
experiment is to observe the titration of
hydrochloric acid a strong acid with
sodium hydroxide a strong base and
acetic acid a

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acid base titration be able to determine
the K_a or K_b from ph data associated
with the titration of a weak acid or base

be able to determine the molar mass of
a solid monoprotic acid from titration
data be able to calculate K_{a1} and K_{a2}
for a polyprotic acid by the end of this
lab students should be able to

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neutralization of hydrochloric acid by
sodium hydroxide the mole ratio of acid
to base is 1 1 $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{O l}$
 $\text{h}_2\text{O l} + \text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{O l}$
one mole of hcl hcl would be fully
neutralized by one mole of naoh naoh
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unknown solid acid the purpose of this
experiment was to determine the
standardization of sodium hydroxide
naoh and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and
the molar mass of an unknown solid
acid using the molar mass of the
standardized sodium hydroxide solution
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 can be used to figure out the
 concentration of an acid or base
 whether an unknown acid or base is
 strong or weak pka of an unknown acid
 or pkb of the unknown base key terms
 titration a process where a solution of
 known strength is added to a certain
 volume of a treated sample containing an

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web conclusion this lab evaluated an
 acid base reaction with the volumetric
 analysis of a standard solution being
 naoh reaction with a 2nd substance
 $\text{h}_2\text{c}_2\text{o}_4 \text{ aq} + 2\text{naoh} \text{ aq} \rightarrow \text{na}_2\text{c}_2\text{o}_4 + 4\text{h}_2\text{o}$ to
 differentiate an acid from a base in a
 solution a ph scale is used the ph scale
 ranges from 0 to 14 a ph of 7 is neutral
 in an acid base reaction

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 titration to explore the concepts of
 stoichiometry and equivalence points
 we found the concentration of an
 unknown substance by mixing 2m hcl
 with the naoh of unknown
 concentration in order to
 experimentally ascertain the
 concentration of the naoh

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web this point of titration is referred to
 as the equivalence point considering
 the 1 1 stoichiometry of this acid base
 reaction $\text{naoh} \text{ aq} + \text{c}_6\text{h}_4\text{cooh} \text{ cook} \text{ aq} \rightarrow$
 $\text{c}_6\text{h}_4\text{coona} \text{ cook} \text{ aq} + \text{h}_2\text{o} \text{ l}$ the point of
 equivalence is the point of titration
 when the number of moles of naoh na
 added is equal to the number of moles
 of khp nb in the