

The Black Hand Book

Albert A Gayle

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pair with the acid and base differing by
the presence or absence of a proton a
conjugate acid base pair for instance
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 H_2CO_3 and its conjugate base the
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an acid and a base react not to produce
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bases disassociate almost or entirely to
completion in aqueous solution

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recognition that some lewis acids and
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nature as salts of chloride or as oxide
ores while others are found in
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common substances found in many
every day items from fruit juices and

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differences between the three

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and bases in living things provide the

pH that enzymes need enzymes are

biological catalysts that must work

effectively for biochemical reactions to

occur most enzymes can do their job

only at a certain level of acidity cells

secrete acids and bases to maintain the

proper pH for enzymes to do their work

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acid and base acids and bases

definition an acid is any hydrogen

containing substance that is capable of

donating a proton hydrogen ion to

another substance a base is a molecule

or ion able to accept a hydrogen ion

from an acid acidic substances are

usually identified by their sour taste

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a binary acid is an acid that consists of

hydrogen and one other element the

most common binary acids contain a

halogen the acid name begins with the

prefix hydro followed by the base name

of the anion followed by the suffix ic

difference between acid and base

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it has a sour taste they produce

hydrogen after reacting with active

metals like zinc magnesium aluminium

or iron acids have a low pH value i e

below 7 salts and water are produced

when an acid reacts with base or alkalis

most of the acids are solids

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important enough that we devote an

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aqueous solutions will conduct

electricity because they contain

dissolved ions therefore acids and

bases are electrolytes strong acids and

bases will be strong electrolytes weak

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electrolytes

6 1 what is an acid and a base

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arrhenius s definition of acids and

bases the earliest definition of acids

and bases is arrhenius s definition

which states that an acid is a substance

that forms hydrogen ions h⁺ when

dissolved in water and a base is a

substance that forms hydroxide ions oh⁻

when dissolved in water

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when equal moles of an acid and a base

are combined the acid is neutralized by the base the products of this reaction are an ionic compound which is labeled as a salt and water

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