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 $2 \text{ICl} + \text{H}_2 \rightarrow 2 \text{HI} + 2 \text{HCl}$   
rate  $k [\text{ICl}]^2 [\text{H}_2]$   
a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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b order of the reaction with respect to acetaldehyde is 3/2 and overall order is also 3/2

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