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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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web chemistry 120 hess s law worksheet calculate Δh for the reaction
 $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ from the following data
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285.8 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ}$

$\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) \quad \Delta h = 285 \text{ kJ}$

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 hess s law example 12.1 determine Δh for the target reaction
 $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ given the following information

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 libretexts *quot hess s law quot* 2019 which determines the enthalpy change value
 hess s law is written as $\Delta h_{\text{target}} = \Delta h_1 + \Delta h_2 + \Delta h_3$ or $\Delta h_{\text{target}} = \sum \Delta h_{\text{known}}$

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two principles must be understood one if an equation is reversed the sign of the Δh value is also reversed two if an equation is multiplied by a

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 $\text{Mg}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{MgO}(\text{s})$ we used *additivity of heats of reaction hess s law lab report - Aug 03 2022*

web hess s law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are
 cohen 2016 the equations for the reactions in the experiment done are as follows
 $\text{NaOH}(\text{s}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{NaOH}(\text{aq})$

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experiment was designed using hess's law and simple calorimetry to determine the enthalpy for naoh's hess's law says that the overall change in enthalpy for the reaction is the sum of all changes regardless of the different stages or phases of a reaction

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$O_2(g) + 2CO_2(g) + 2H_2O(l) \rightarrow H_2C_2H_4(g) + 3\frac{1}{2}O_2(g) + 2CO_2(g) + 3H_2O(l)$ $\Delta H = 1411 \text{ kJ}$ $C_2H_6(g) + \frac{3}{2}O_2(g) \rightarrow 2CO_2(g) + 3H_2O(l)$ $\Delta H = 1560 \text{ kJ}$

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web thus for reaction 9.2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change ΔH for products $\Delta H_{\text{products}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}}$

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