

S Beginning With I

Rachel S Tattersall

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solution is 1 0 25m 2 0 5m 3 0 888m 4 none of
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which is more important molality or molarity both molality
and molarity are concentration units and they are used
interchangeably

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web molarity is the number of moles of a substance per litre
 of solution also known as molar concentration a capital M
 signifies solutions labelled with molar concentration a 1.0 M
 solution contains 1 mole of solute per litre of solution
 molality is the number of solute moles per kilogram
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9579 g mol x 0 0129 m 2 molality 0 825 g 141 9579 g mol 0
00581158 mol 0 00581158 mol 0 4500 kg 0 0129 m 3 mole
fraction na 2 hpo 4 0 825 g 141 9579 g mol 0 00581158 mol
h 2 o 450 0 g 18 015 g mol 24 97918401 mol

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4 2 2 a first convert 250 ml to liters 250 1000 0 25 then
calculate molarity 5 moles 0 25 liters 20 m 3 c a solution
with molarity 2 requires 2 m of n a oh per liter so 4 x 2 8 m 4
a a solution of molarity 1 5 m requires 1 5 mol of na to every
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of a mixture mmix can be calculated using the following formula $m_{\text{mix}} = \frac{m_1 v_1 + m_2 v_2}{v_1 + v_2}$ $\frac{1.5 \times 480 + 1.2 \times 520}{480 + 520} = \frac{720 + 624}{1000} = 1.344$ m q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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