

# Quotes From The 48 Laws Of Power

A Loxley

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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example  $\text{hcl naoh nacl h}_2\text{o}$   
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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web concerning alkali and acid being an answer to a letter to a physician concerning acid and alkali to which is added a specimen of a new hypothesis etc jan 08 2021 acids alkalis and salts classic reprint sep 15 2021 excerpt from acids alkalis and salts it has often been said and still more often implied that

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kj less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with hcl as 57 9 kj mol 1

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web sep 20 2014 i will illustrate this with some examples

when an acid is neutralised by an alkali the reaction is exothermic eg 1  $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$  for which  $\Delta H = -57 \text{ kJ mol}^{-1}$  eg 2  $\text{HNO}_3(\text{aq}) + \text{KOH(aq)} \rightarrow \text{KNO}_3(\text{aq}) + \text{H}_2\text{O(l)}$  for which  $\Delta H = -57 \text{ kJ mol}^{-1}$

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example  $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$  you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts