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3 4 5 br almost

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equations are useful to show which  
chemicals react together reactants and  
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how many have in common they are  
questions about the amount of a  
substance and are similar to questions  
scientists ask 2 list two or

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calculate the number of moles from  
each of the following a  $6.95 \times 10^{23}$   
atoms of cu b  $7.12 \times 10^{21}$  molecules of  
co c  $3.33 \times 10^{22}$  ions of pb d  $7.00 \times$   
 $10^{23}$  formula units of lino 3

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quantities of reactants or products in a  
chemical reaction using the  
relationships found in the balanced  
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the quantity of heat for a process is  
represented by the letter q the sign of q  
for an endothermic process is positive  
because the system is gaining heat a

chemical reaction or physical

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one milligram is equal to 0 001 grams  
790 drag each number to the correct location sort the numbers based on the number of significant figures they have one significant figure 1 000

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the quantities commonly encountered in chemical problems include the number of moles of a substance the number of atoms molecules or formula units of a substance and the mass in

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it is always a good idea to estimate the answer before you do the actual calculation in this case the mass given 35 00 g is less than the molar mass so the answer should be less than 1

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the atomic mass of an element expressed in grams is the mass of a mole of the element the mass of a mole of an element is the molar mass to calculate the molar mass of a compound

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