

# La Plaa Del Diamante Gratis

Peipei Pang

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 calorimeter th water in it and the hot  
 object made up of copper which is put  
 in the calorimeter are the same the  
 initial temperature of the calorimeter  
 and water is 30 c and that of hot object  
 is 60 c the specific heat of copper and  
 water are 0 09 cal gm c and 1 cal gm c  
 respectively

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 equation  $q = \text{specific heat} \times m \times \Delta t$  where

$q$  is heat flow  $m$  is mass in grams and  
 $\Delta t$  is the temperature change plugging  
 in the values given in the problem you  
 get  $q_{\text{water}} = 418 \text{ J/g} \times 110 \text{ g} \times 26.6 \text{ }^\circ\text{C}$   
 $25.0 \text{ }^\circ\text{C}$   $q_{\text{water}} = 550 \text{ J}$   $\Delta h$   $q_{\text{water}} = 550 \text{ J}$   
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 problems  $q_{\text{sur}} = m \times c \times \Delta t$   $q_{\text{heat}} = m \times \text{mass}$   
 $q_{\text{rxn}} = q_{\text{sur}} = \Delta t$  name per date c

specific heat for water 4184 J/gC  
 what is the specific heat of aluminum if  
 the temperature of a 28.4 g sample of  
 aluminum is increased by 8.1 °C when  
 207 J of heat is added 2

### 7 3 heats of reactions and

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 the heart of all calorimetry problems  
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 released or absorbed at constant  
 pressure is equal to  $\Delta H$  the relationship  
 between heat and  $\Delta H_{\text{rxn}}$  is  $\Delta H_{\text{rxn}} = q_{\text{calorimeter}}$   
 $m \times c \times \Delta t$   $7.314 \text{ J/g} \times 14.7 \text{ g} \times \Delta t$   
 $\Delta H_{\text{rxn}} = q_{\text{calorimeter}} = m \times c \times \Delta t$   
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have 125 grams of water at a  
 temperature of 25.0 °C and heat it to  
 145.0 °C what will the change in enthalpy  
 be 2 sketch the heating curve  
 temperature vs time for the process in  
 1.3 i have 18.7 grams of an unknown  
 liquid chemical at a temperature of 35.0  
 °C

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 ml of 1.00 M HCl at 25.0 °C is added to  
 25.0 ml of 1.00 M NaOH at 25.0 °C in a  
 foam cup calorimeter a reaction occurs  
 and the temperature rises to 32.0 °C  
 calculate the enthalpy change in kJ for  
 this reaction assume the densities of  
 the solutions are 1.00 g/ml and that  
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 consider a simple example that  
 illustrates the core idea behind  
 calorimetry suppose we initially have a  
 high temperature substance such as a  
 hot piece of metal  $m$  and a low  
 temperature substance such as cool  
 water  $w$  answer the initial temperature

of the copper

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an apparatus used for calculating the

heat developed during a chemical

mechanical or electrical reaction it also

helps to measure the heat capacity of

various

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bomb calorimeter with a total heat

capacity of 15 48 kj c the temperature

of the calorimeter and its contents

changes from 19 77 c to 33 90 c

calculate e and h for the following

reaction as written 2 nh 3 g 3 o 2 g 2 ch

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temperature is closer to t 1 than to t 2 a

the final temperature is closer to t 1

than to t 2 the final temperature is

exactly halfway between t 1 and t 2 b

the final temperature is exactly halfway

between t 1 and t 2 the final

temperature is closer to t 2 than to t 1 c

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will go over the main types of questions

on calorimetry including the heat

capacity the heat of reaction finding the

final temperature of a mixture constant

pressure calorimetry and constant

volume calorimetry a few important

concepts and formulas you will need to

solve these questions

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moles of hexane c6h14 is combusted in

a bomb calorimeter containing 5 65

liters of water calculate the molar heat

of combustion of hexane if the water

temperature rises 55 4 c the specific

heat capacity of water is 4 184 j g c h

ms t h 5 650 grams h2o 4 184 j g c 55 4

c h 1310 kj

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web 6 a 218 g sample of steam at 121oc

is cooled to ice at 14oc find the change

in heat content of the system 7 in going

from ice at 34oc to steam at 138oc a

sample of water absorbs 1 41 x 105 j

find the mass of the sample answers 1 2

68 x 104 j 2 6 23 x 104 j 3 1 71 x 104 j 4

1 11 x 105 j 5 2 11 x 105 j 6

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