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**Rachel Sandford**

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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton  $h_a + b \rightleftharpoons h_b$  strong acids e g sulfuric acid have a very pronounced

tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from ph data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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at the beginning of

the titration the solution in the erlenmeyer flask is acidic as the base is added it completely reacts with the acid and the solution in the erlenmeyer flask continues to be acidic but at the equivalence point

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react with a given amount of another substance in this activity your goal is to determine the molar concentration and strength of two acid solutions by conducting titrations with a base solution of known

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web 1 approximately 3 g of hydrated oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the

oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then  $\text{ph} = \log 0.100 = 1$

000  $\text{ph} = \log 0.100 = 1$  000 b titrant volume 12.50 ml

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sodium hydroxide when the acid and base react they form nacl sodium chloride which is also known as table salt

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