

# Gold Experience C1 Answer Key

**M Lipman**

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members and their behind the scenes and her own role as a founding member of wings gave her yet another point of view on musical st

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web produced in close collaboration with her family linda mccartney life in photographs gathers some of the finest pictures from mccartney s astonishing archive of over 200 000 images from the epicenter of 1960s rock to rural bliss with her husband and kids mccartney s lens captured fame family music animals and the english countryside

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web jun 1 2011 life in photographs is an excellent and long overdue retrospective of linda mccartney s entire photographic oeuvre the early part of the book features an interesting variety of musicians rare beatles photographs from the white album sessions are included and a backstage shot of a yawning jimi hendrix is particularly fantastic

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web mcartney s photographs are true to how she experienced life from the extraordinary photos of her time in tucson to pictures of the dynamic music scene of the 1960s to images of her domestic life in london in 1968 she became the first female photographer to have her work featured on the cover of rolling stone with her photo of eric clapton

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$$\begin{aligned} & \text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3 \\ & \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2 \\ & 2 \text{NaCl} + \text{F}_2 \rightarrow 2 \text{NaF} + \text{Cl}_2 \\ & 2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O} \\ & \text{Pb} + 2 \text{HCl} \rightarrow \text{H}_2 + \text{PbCl}_2 \\ & 2 \text{Al} + 3 \text{K}_2\text{SO}_4 \rightarrow 6 \text{KBr} + \text{Al}_2\text{SO}_4 + 3 \text{CH}_4 \\ & 2 \text{O}_2 \rightarrow 1 \text{CO}_2 + 2 \text{H}_2\text{O} \\ & \text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O} \\ & \text{C}_8\text{H}_{18} + 25 \text{O}_2 \rightarrow 16 \text{CO}_2 + 18 \text{H}_2\text{O} \end{aligned}$$

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gaseous butane c 4 h 10 reacts with diatomic oxygen gas to yield gaseous carbon dioxide and water vapor

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2 1 be no 3 2 2 rbf

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 equations below

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$$2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$$

$$2 \text{ NaCl} + \text{F}_2 \rightarrow 2 \text{ NaF} + \text{Cl}_2$$

$$\text{Cl}_2 + 4 \text{ H}_2\text{O} + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}_5 + 1 \text{ Pb(OH)}_2$$

$$2 \text{ HCl} + 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2 \rightarrow 6 \text{ H}_2\text{O} + 3 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4$$

$$6 \text{ KBr} + 1 \text{ Al}_2\text{SO}_4 \rightarrow 3 \text{ H}_2\text{O} + 7 \text{ CH}_4 + 2 \text{ O}_2 + 1 \text{ CO}_2$$

$$2 \text{ H}_2\text{O} + 8 \text{ C} + 3 \text{ H}_2\text{SO}_4 \rightarrow 2 \text{ CO}_2 + 4 \text{ H}_2\text{O}$$

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 $\text{O}_2$  2  $\text{NaCl}$  1  $\text{F}_2$   $\rightarrow$  2  $\text{NaF}$  1  $\text{Cl}_2$  2  $\text{H}_2$  1  $\text{O}_2$   $\rightarrow$   
 2  $\text{H}_2\text{O}$   $\text{Pb}$   $\text{OH}$  2 2  $\text{HCl}$   $\rightarrow$  2  $\text{H}_2\text{O}$  1  $\text{PbCl}_2$   
 $\text{AlBr}_3$  3  $\text{K}_2\text{SO}_4$   $\rightarrow$  6  $\text{KBr}$  1  $\text{Al}_2\text{SO}_4$  3  $\text{CH}_4$  2  
 $\text{O}_2$   $\rightarrow$  1  $\text{CO}_2$  2  $\text{H}_2\text{O}$   
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 equations coefficients equal to one 1 do  
 not need to be shown in your answers 2  
 $\text{Fe}$  3  $\text{Cl}_2$  2  $\text{FeCl}_3$  4  $\text{Fe}$  3  $\text{O}_2$  2  $\text{Fe}_2\text{O}_3$  2  $\text{FeBr}_3$   
 3 3  $\text{H}_2\text{SO}_4$  1  $\text{Fe}$  2  $\text{SO}_4$  3  $\text{d}$  1  $\text{C}_4\text{H}_{10}$  6  $\text{O}_3$  1  $\text{H}_2\text{O}$  2  $\text{C}_2\text{H}_4\text{O}_2$   
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 gcse answers chemistrytutor me 2018  
 page 3 of 3 38  $\text{Fe}$  2  $\text{O}$  3 3 $\text{CO}$  2 $\text{Fe}$  3 $\text{CO}$  2 39  
 2 $\text{H}$  3  $\text{PO}_4$  3 $\text{BaO}$   $\text{Ba}$  3  $\text{PO}_4$  2 3 $\text{H}_2$  2  $\text{O}$  40  
 2 $\text{MgNH}_4$  4  $\text{PO}_4$  4  $\text{Mg}$  2  $\text{P}$  2  $\text{O}$  7 2 $\text{NH}_3$  3  $\text{H}_2$   $\text{O}$   
 41 2 $\text{Al}$  3 $\text{H}_2$  2  $\text{SO}_4$   $\text{Al}$  2  $\text{SO}_4$  3 3 $\text{H}_2$  2 42  $\text{P}$  4  
 $\text{O}$  10 3 $\text{HCl}$   $\text{POCl}_3$  3  $\text{HPO}_3$  3 43 4 $\text{NO}$  2 $\text{H}_2$   $\text{O}$   
 3 $\text{O}$  2 4 $\text{HNO}_3$   
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 equation described in section 4 1 is  
 balanced meaning that equal numbers  
 of atoms for each element involved in  
 the reaction are represented on the  
 reactant and product sides this is a

requirement the equation must satisfy  
 to be consistent with the law of  
 conservation of matter  
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 chemical equations are symbolic  
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 substances undergoing the change  
 reactants and substances generated by  
 the change products are  
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 equation gives the number and type of  
 atoms participating in a reaction the  
 reactants products and direction of the  
 reaction balancing an unbalanced  
 equation is mostly a matter of making  
 certain mass and charge are balanced  
 on the reactants and products side of  
 the reaction arrow  
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 when a chemical equation is balanced it  
 means that equal numbers of atoms for

each element involved in the reaction are represented on the reactant and product sides this is a requirement the equation must satisfy to be consistent with the law of conservation of matter

### **balancing chemical equations**

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web mar 14 2023 the key to balancing chemical equations the ultimate goal for balancing chemical equations is to make both sides of the reaction the reactants and the products equal in the

number of atoms per element this stems from the universal law of the conservation of mass which states that matter can neither be created nor destroyed

### **answer key for the balance**

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balance the following chemical equation  
 $\text{mg} + \text{oh} + 2 \text{hcl} + \text{mgcl}_2 + \text{h}_2 + \text{o}_2$   
note all reactants and products require a coefficient of at least one  
stuck review related articles videos or use a hint  
report a problem do 4 problems